

**Directions:** Use website homework page to help complete this assignment.

## Vocabulary Practice

*Study 5-10 minutes a Night*

I.N. p.8

Text book L: Chemical Interactions

<u>In a Word</u> (Write one word that is not the key term that captures what the word means)	<u>In a Sentence</u> (Write a sentence using the word in correct context.)	<u>In a Picture</u> (Draw a detailed picture with labels and purposeful coloring to help you understand)	<u>Key Terms</u> (Study and practice words 5-10 minutes a day outside of class.)
Physical, mental, and other representations of an idea to help people understand a concept that they cannot observe. (p.8) _____			<b>model</b>
The smallest particle of an of any molecule. (p.7) _____			<b>atom</b>
Two or more atoms bonded together. (p.31) _____			<b>molecule</b>
The study of properties of matter how matter changes. (p.46) _____			<b>chemistry</b>
Anything that has mass and occupies space. (p.6,46) _____			<b>matter</b>
Two or more substances that are mixed together but not chemically combined (p.7). _____			<b>mixture</b>

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## Vocabulary Practice

*Study 5-10 minutes a Night*

I.N. p.9

Text book L: Chemical Interactions

<u>In a Word</u> (Write one word that is not the key term that captures what the word means)	<u>In a Sentence</u> (Write a sentence using the word in correct context.)	<u>In a Picture</u> (Draw a detailed picture with labels and purposeful coloring to help you understand)	<u>Key Terms</u> (Study and practice words 5-10 minutes a day outside of class.)
The force that holds atoms together. (p.13)  _____			<b>chemical bond</b>
A substance that causes a chemical reaction. (p.57)  _____			<b>reactants</b>
A substance formed due to a chemical reaction. (p.57)  _____			<b>products</b>
The principle stating that matter is not created nor destroyed during a chemical reaction. (p.58)  _____			<b>law of conservation of mass</b>
A reaction that absorbs energy in the form of heat (p.52)  _____			<b>endothermic reaction</b>
A reaction that releases energy in the form of heat. (p.53)  _____			<b>exothermic reaction</b>